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ph and poh continued 1. 0.01 m hcl $[H^+] = 0.01\text{M}$ $\text{pH} = -\log(0.01) = 2$ 2. 0.0010 m naoh ... $\text{pOH} = -\log(0.150) = 0.8$
→ $\text{pH} = 13.2$ 6. 2.0 m H_2SO_4 (assume 5.0% ...)

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